

PERIODIC TABLE WORKSHEET

Name _____

1. Where are the most active metals located? _____
2. Where are the most active nonmetals located? _____
3. As you go from left to right across a period, the atomic size (decreases / increases). Why? _____
4. As you travel down a group, the atomic size (decreases / increases). Why? _____
5. A negative ion is (larger / smaller) than its parent atom.
6. A positive ion is (larger / smaller) than its parent atom.
7. As you go from left to right across a period, the first ionization energy generally (decreases / increases). Why? _____
8. As you go down a group, the first ionization energy generally (decreases / increases). Why? _____
9. Where is the highest electronegativity found? _____
10. Where is the lowest electronegativity found? _____
11. Elements of Group 1 are called _____
12. Elements of Group 2 are called _____
13. Elements of Group 3-12 are called _____
14. As you go from left to right across the periodic table, the elements go from (metals / nonmetals) to (metals / nonmetals).
15. Group 17 elements are called _____
16. The most active element in Group 17 is _____
17. Group 18 elements are called _____
18. What sublevels are filling across the Transition Elements? _____
19. Elements within a group have a similar number of _____
20. Elements across a series have the same number of _____
21. A colored ion generally indicates a _____
22. As you go down a group, the elements generally become (more / less) metallic.
23. The majority of elements in the periodic table are (metals / nonmetals).
24. Elements in the periodic table are arranged according to their _____
25. An element with both metallic and nonmetallic properties is called a _____