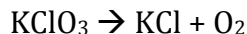
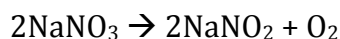


Limiting Reagent Practice Problems

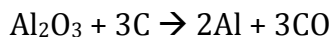
1. 3.45 moles of Nitrogen gas reacts with 4.85 moles of Hydrogen gas to form Ammonia (NH₃). What is the limiting reactant? How many moles of Ammonia will form?
2. Calculate the number of grams of Potassium Chloride that will be formed by the decomposition of 6.45 g of Potassium Chlorate.



3. If 156 g of Sodium Nitrate react and 112 g of Sodium Nitrite are recovered, what is the percentage yield?



4. How many grams of Aluminum form from 9.73 g of Aluminum oxide if the yield is 91%?



5. 25.3 g of Magnesium reacts with 44.3 g of Copper(II)Nitrate to form Copper and Magnesium Nitrate. What will be the limiting and excess reagents? Determine the percentage yield if the actual yield is 14.3 g of Copper.
6. How many grams of Aluminum Chloride could be produced from 34.0 g of Aluminum and 39.0 g of Chlorine gas? How many grams of the excess reagent will be left over?