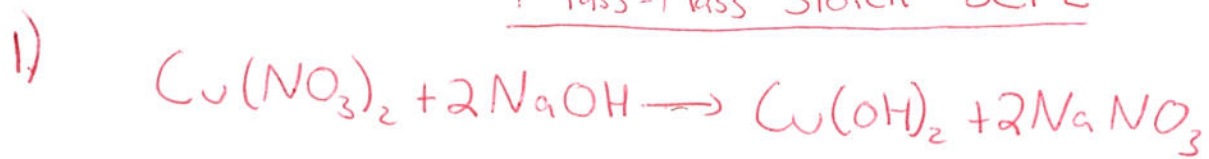
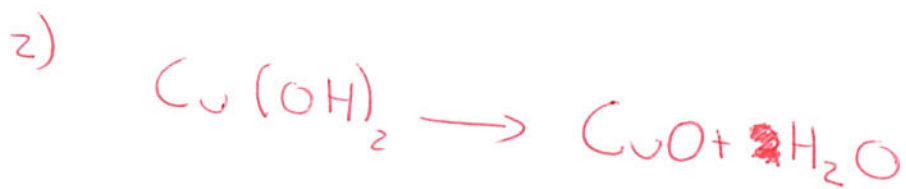


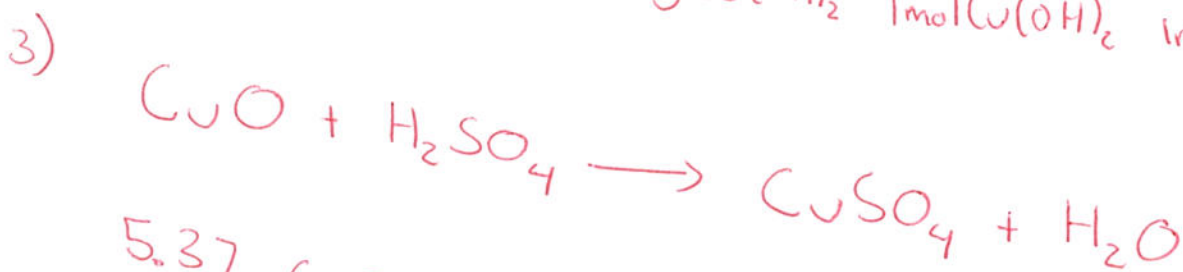
Mass-Mass Stoich - Set 2



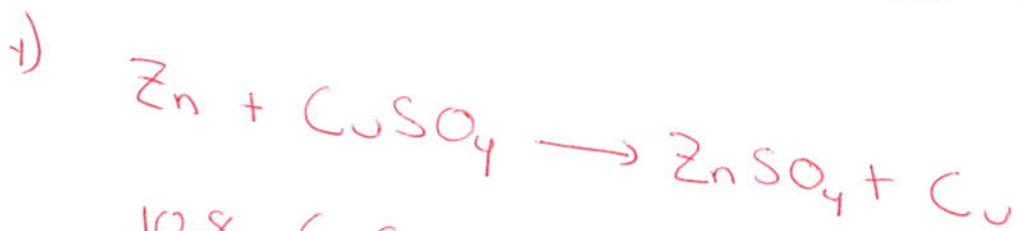
$$12.7 \text{ g Cu}(\text{NO}_3)_2 \times \frac{1 \text{ mol Cu}(\text{NO}_3)_2}{187.5 \text{ g Cu}(\text{NO}_3)_2} \times \frac{1 \text{ mol Cu}(\text{OH})_2}{1 \text{ mol Cu}(\text{NO}_3)_2} \times \frac{97.5 \text{ g Cu}(\text{OH})_2}{1 \text{ mol Cu}(\text{OH})_2} = \underline{\underline{6.60 \text{ g}}}$$



$$6.59 \text{ g Cu}(\text{OH})_2 \times \frac{1 \text{ mol Cu}(\text{OH})_2}{97.5 \text{ g Cu}(\text{OH})_2} \times \frac{1 \text{ mol CuO}}{1 \text{ mol Cu}(\text{OH})_2} \times \frac{79.5 \text{ g CuO}}{1 \text{ mol CuO}} = \underline{\underline{5.37 \text{ g}}}$$



$$5.37 \text{ g CuO} \times \frac{1 \text{ mol CuO}}{79.5 \text{ g CuO}} \times \frac{1 \text{ mol CuSO}_4}{1 \text{ mol CuO}} \times \frac{159.6 \text{ g CuSO}_4}{1 \text{ mol CuSO}_4} = \underline{\underline{10.8 \text{ g CuSO}_4}}$$



$$10.8 \text{ g CuSO}_4 \times \frac{1 \text{ mol CuSO}_4}{159.6 \text{ g CuSO}_4} \times \frac{1 \text{ mol Cu}}{1 \text{ mol CuSO}_4} \times \frac{63.5 \text{ g Cu}}{1 \text{ mol Cu}} = 4.30 \text{ g Cu}$$

5) They're the same!