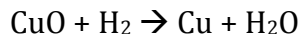


Stoichiometry Review

1. How many grams of silver chloride are produced from 5.0 g of silver nitrate reacting with an excess of barium chloride?
2. Use the equation below. What is the limiting reactant when 19.9 g CuO react with 2.02 g H₂. The actual yield of copper was 15.0 g. What is the percentage yield?



From the equation above, how many grams of Cu can be collected if 20.6 g CuO react with an excess of hydrogen with a yield of 91%?

3. When Cu(OH)₂ is heated, it decomposes to CuO and H₂O. How many grams of H₂O will be formed from the decomposition of 6.59 g of Cu(OH)₂?
4. If 20 g of Lithium Hydroxide react with an excess of Potassium Chloride, what is the theoretical yield of Lithium Chloride? If I actually produce 6 g of Lithium Chloride, what is my percent yield?
5. If I start with 45 g of C₂H₄, how many grams of carbon dioxide will be produced?



6. Silver Nitrate and Sodium Phosphate are reacted in equal amounts of 200.0 g each. How many grams of silver phosphate are produced? What are the limiting and excess reactants? How much of the excess reactant is leftover after the reaction has gone to completion?